

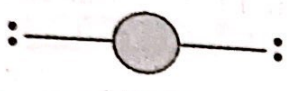
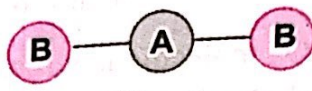
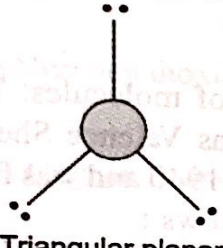
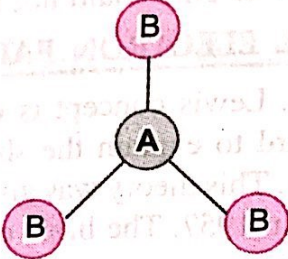
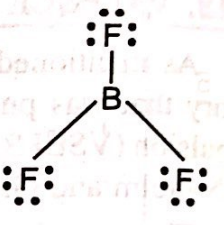
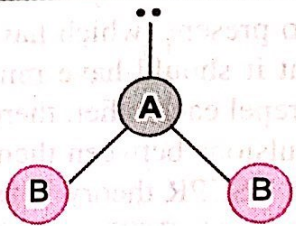
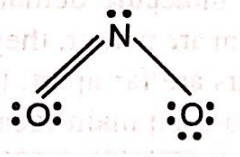
Total No. of electron pairs	Geometry of the electron pairs	Bond pairs	Lone pairs	Geometry (Shape) of the molecule	Illustrative Example (Other Examples)
		(Molecular Formula)			
2	 Linear	2	0	 Linear	$\ddot{\text{O}} = \text{C} = \ddot{\text{O}}$ (BeF ₂ , BeCl ₂ , HgCl ₂)
3	 Triangular planar	3	0	 Triangular planar	 (AlCl ₃ , SO ₃)
		(AB ₃)			
		2	1	 Bent (V-shape)	 (SO ₂ , O ₃ , SnCl ₂)
		(AB ₂ L)			

Table Continued...

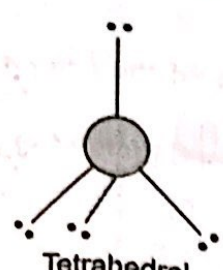
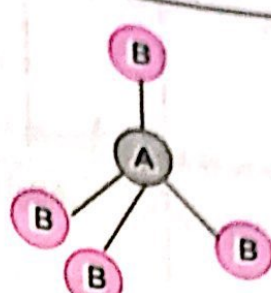
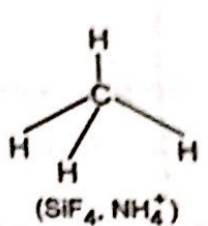
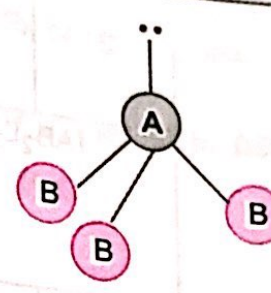
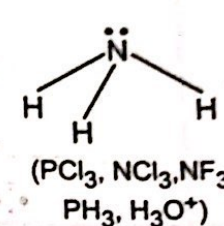
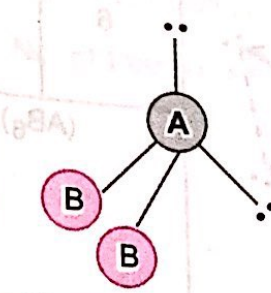
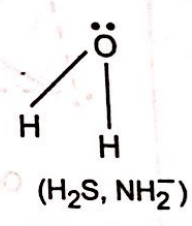
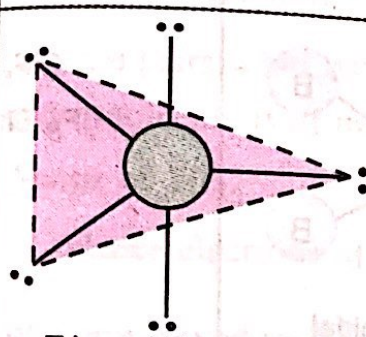
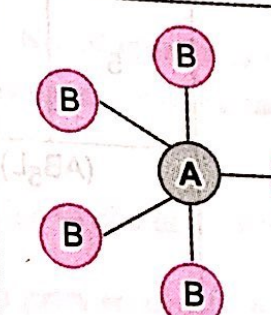
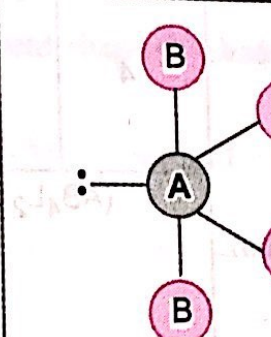
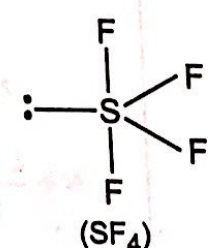
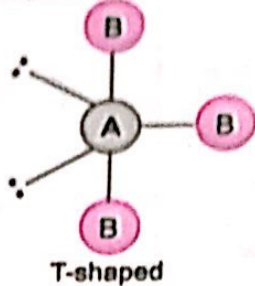
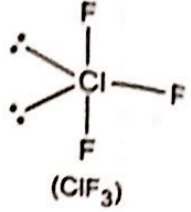
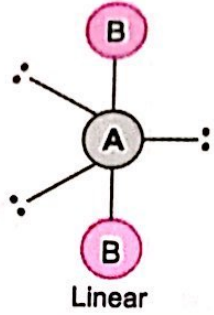
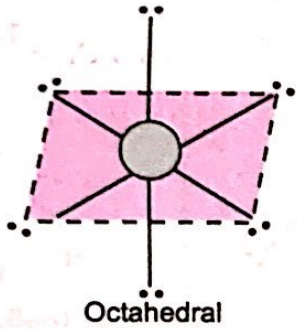
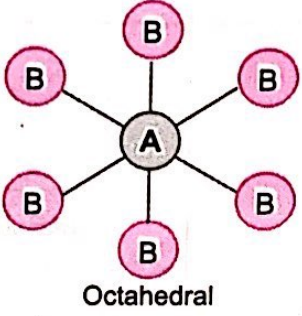
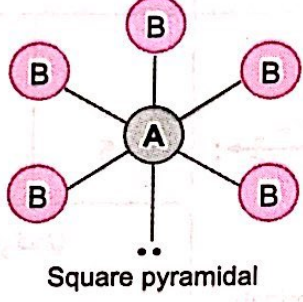
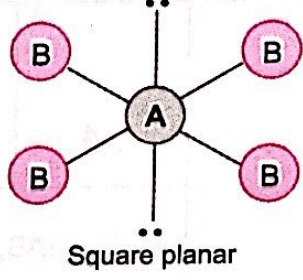
<p>4</p>  <p>Tetrahedral</p>	<p>4</p> <p>0</p> <p>(AB₄)</p>	 <p>Tetrahedral</p>	 <p>(SiF₄, NH₄⁺)</p>
	<p>3</p> <p>1</p> <p>(AB₃L)</p>	 <p>Trigonal pyramidal</p>	 <p>(PCl₃, NCl₃, NF₃, PH₃, H₃O⁺)</p>
	<p>2</p> <p>2</p> <p>(AB₂L₂)</p>	 <p>Bent</p>	 <p>(H₂S, NH₂⁻)</p>
<p>5</p>  <p>Trigonal bipyramidal</p>	<p>5</p> <p>0</p> <p>(AB₅)</p>	 <p>Trigonal bipyramidal</p>	<p>PCl₅ (PF₅, SbCl₅, AsF₅)</p>
	<p>4</p> <p>1</p> <p>(AB₄L)</p>	 <p>See saw</p>	 <p>(SF₄)</p>

Table Continued...

		3	2	 <p>T-shaped</p>	 <p>(ClF₃)</p>
		(AB ₃ L ₂)			
		2	3	 <p>Linear</p>	<p>XeF₂ (I₃⁻, Br₃⁻, ICl₂⁻)</p>
		(AB ₂ L ₃)			
6	 <p>Octahedral</p>	6	0	 <p>Octahedral</p>	<p>SF₆</p>
		(AB ₆)			
		5	1	 <p>Square pyramidal</p>	<p>ClF₅ (IF₅, BrF₅)</p>
		(AB ₅ L)			
		4	2	 <p>Square planar</p>	<p>XeF₄</p>
		(AB ₄ L ₂)			